

## Types of Chemical Reactions

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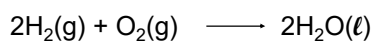
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### Synthesis Reactions

- Two or more reactants combine to produce a new product.



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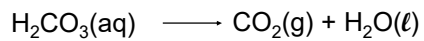
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### Decomposition Reactions

- A compound breaks down into two or more simpler compounds or elements.



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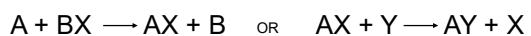
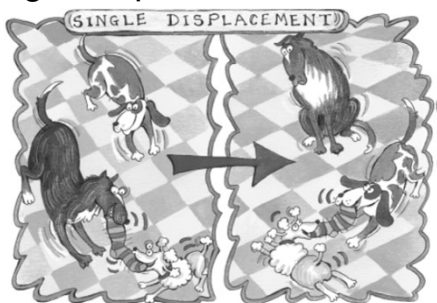
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## Single Displacement Reactions



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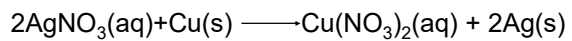
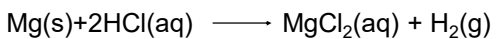
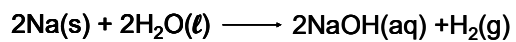
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## Single Replacement Reactions

- One element takes the place of (displaces) another element in a compound.



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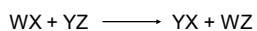
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## Double Displacement Reactions



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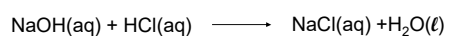
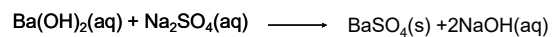
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## Double Replacement Reactions

- The cations of two different compounds exchange places, forming two new compounds.



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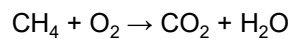
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## Combustion Reaction

- Reaction that occurs when something burns.
- Shown by having oxygen,  $\text{O}_2$ , as one of the reactants.
- The products will **always** be carbon dioxide,  $\text{CO}_2$  and water,  $\text{H}_2\text{O}$



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